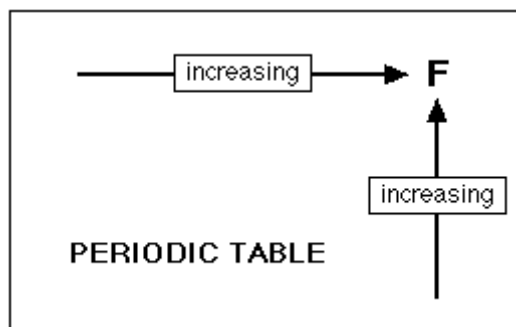


Chemguide – questions

ELECTRONEGATIVITY

You may need a copy of the Periodic Table.

Electronegativity varies around the Periodic Table (ignoring the noble gases) in the following way:



Here are some values you will need to answer question 2:

H 2.1

C 2.5 N 3.0 O 3.5 F 4.0

Cl 3.0

Br 2.8

- What do you understand by the term electronegativity?
 - Explain why the electronegativity of fluorine is greater than the electronegativity of carbon.
 - Explain why the electronegativity of fluorine is greater than the electronegativity of chlorine.

2. Use the figures above to answer the following questions.

a) List these bonds in order of increasing polarity:

H-F F-F C-Cl C-Br C-O N-H

b) By writing δ^+ and δ^- as appropriate above each of the atoms in the bond, show the polarity of the following bonds:

C-O C-Cl C-Br C-N C-C N-H H-Br O-H